

UNITED STATES PATENT APPLICATION

OF

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FOR

AN ULTRA-THIN INTEGRAL COMPOSITE MEMBRANE

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AN ULTRA-THIN INTEGRAL COMPOSITE MEMBRANE

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BACKGROUND OF THE INVENTION

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U.S. Patent No. 4,604,170 to Miyake et al. relates to a multi-layered diaphragm for electrolysis comprising a porous layer of a fluorine-containing polymer having a thickness of from 0.03 to 0.4 mm with its interior and anode-side surface being hydrophilic and an ion exchange layer on its cathode surface with the ion exchange layer being thinner than the porous layer but of at least 0.005 mm and the total thickness of the diaphragm being from 0.035 to 0.50 mm.

20 U.S. Patent 4,865,925 to Ludwig, et al. relates to a gas permeable electrode for electrochemical systems. The electrode includes a membrane located between and in contact with an anode and a cathode. The membrane, which may be made of expanded polytetrafluoroethylene, may be treated with an ion exchange membrane material with the resulting membrane
25 maintaining its permeability to gas. Membrane thicknesses are described to be between 1 and 10 mils, (0.025-0.25 mm), with thicknesses of less than 5 mils (0.125) to be desirable. Examples show that membrane thicknesses range from 15 to 21 mils.

Japanese Patent Application No. 62-240627 relates to a coated or an
30 impregnated membrane formed with a perfluoro type ion exchange resin and
porous PTFE film to form an integral unit. No water or surfactant were used in
the manufacture of this membrane. The combination is accomplished by
fusion or by coating and does not provide for permanent adhesion of the ion
exchange resin to the inside surface of the PTFE film. The weight ratio of the
35 perfluoro ion exchange resin to PTFE is described to be in the range of 3 to
90% with a preferable weight ratio of 10 to 30%.

Japanese Application No. 62-280230 and 62-280231 relate to a composite structure in which a scrim or open fabric is heat laminated and

encapsulated between a continuous perfluoro ion exchange membrane and an ePTFE sheet thus imparting tear strength to the structure. The composite membrane was not used for ionic conduction.

Additional research has also been conducted on the use of perfluorosulfonic acid polymers with membranes of expanded porous polytetrafluoroethylene such as that described in Journal Electrochem. Soc., Vol. 132, No. 2, February 1985, p. 514-515. The perfluoro ion exchange material was in an ethanol based solvent without the presence of water or surfactant. Moreover, ultrasonic energy in the treatment of this membrane.

Heretofore and as represented by the references discussed above, there is a need for an integral ultra-thin strong ion exchange composite membrane, with long term chemical and mechanical stability that has a thickness before swelling of at most 1 mil (0.025 mm), with more than 90% of the pore volume of the membrane filled with a perfluoro ionomer to render it at least substantially occlusive and that is capable of swelling without deterioration of mechanical properties.

SUMMARY OF THE INVENTION

An ultra-thin integral composite membrane is provided including a porous polymeric membrane having a structure of micropores of polymer with a porosity of greater than 35%, an average pore diameter of less than 10 microns and a thickness of at most 0.025 mm and a perfluoro ion exchange polymer impregnated within the micropores so as to render the micropores substantially occlusive, wherein the composite membrane is impermeable to gases and liquids and is substantially free of pinholes. Porous polymeric membranes suitable for this invention include membranes made of perfluoroalkyloxy resin, fluorinated ethylene propylene, polyolefins, polyamides, cellulotics, polycarbonates, ^{and} fluorinated and chlorinated polymers, ~~and polysulfones~~. Perfluoro ion exchange materials suitable for use with this invention include perfluorinated sulfonic acid resin, perfluorinated carboxylic acid resin, polyvinyl alcohol, divinyl benzene, and styrene based polymers. A reinforcement backing may also be provided.

Methods for making the ultra-thin integral composite membranes are also provided.

BRIEF DESCRIPTION OF THE DRAWINGS

5 Figure 1 is a schematic cross-section of the composite membrane that is fully impregnated with the ion exchange material.

Figure 2 is a schematic cross-section of the composite membrane that is fully impregnated with the ion exchange material and has a backing material attached thereto.

10 Figure 3 is a photomicrograph of a cross-section of expanded PTFE that has not been treated with any ion exchange material at a magnification of 2.5 kX.

Figure 4 is a photomicrograph of a cross-section of expanded PTFE fully impregnated with an ion exchange material at a magnification of 5.1 kX.

15 DETAILED DESCRIPTION OF THE INVENTION

20 An ultra-thin composite membrane is provided and includes a base material of microporous membrane with a thickness less than 1 mil (0.025 mm) having a microstructure of micropores and perfluoro ion exchange resin that substantially impregnates the microporous membrane so as to occlude the micropores. The ultra-thin composite membrane may be employed in many different types of applications including for example, chemical separation, electrolysis in fuel cells and batteries, pervaporation, gas separation, dialysis separation, industrial electrochemistry such as chlor-alkali, and other
25 electrochemical devices, catalysis as a super acid catalyst and use as a *medium of catalyst support* in enzyme immobilization.

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30 The ultra-thin composite membrane is mechanically strong and is substantially and uniformly pore occlusive so that it is particularly useful as an ion exchange material. Ultra-thin is hereby defined as 1 mil (0.025 mm) or less. Uniform is hereby defined as continuous impregnation with the ion exchange material so that no pin holes or other discontinuities exist within the composite structure. In addition, pore occlusive is hereby defined as pores being substantially impregnated (i.e., at least 90%) with the perfluoro ion exchange material rendering the final material air impermeable with a Gurley number of infinity.
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The microporous membrane which serves as the base material for the composite has a porosity of greater than 35% and preferably has a porosity of between 70-95%. The pores of the microstructure have a diameter less than

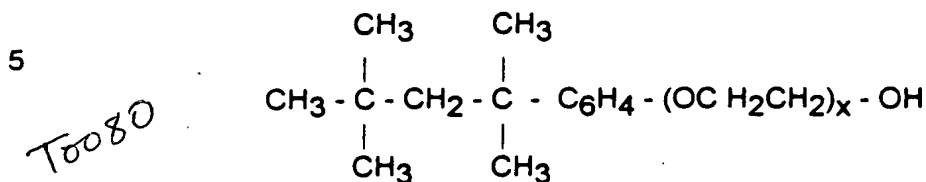
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A most preferred surfactant is a nonionic material, octylphenoxy polyethoxyethanol having a chemical structure:



10 where $x = 10$ (average)

known as Triton® X100, commercially available from Rohm & Haas of Philadelphia, PA.

15 Figure 1 shows a schematic view of the composite membrane with ion exchange material 2 and the base material 4 so that the micropores of the interior of the base material 4 are fully impregnated. The final composite membrane has a uniform thickness free of any discontinuities or pinholes on the surface. The micropores of the membrane are 100% occluded thus causing the composite membrane to be impermeable to liquids and gases.

20 Figure 4 shows a scanning electron photomicrograph of this composite membrane.

Alternatively, the ion exchange material and surfactant mixture 2 may be applied to the membrane 4 so that most of the pores are uniformly treated rendering the membrane substantially impregnated with the ion exchange material. The composite membrane is still free of any discontinuities or pinholes.

Optionally, and as shown schematically in Figure 2, the composite membrane may be reinforced with a woven or non-woven material 6 bonded to one side of the membrane. Suitable woven materials include for example, 30
scrims made of woven fibers of expanded porous polytetrafluoroethylene, commercially available from W. L. Gore & Associates, Inc., of Elkton, Maryland; webs made of extruded or oriented polypropylene netting commercially available from Conwed, Inc. of Minneapolis, MN; and woven materials of polypropylene and polyester of Tetko Inc., of Briarcliff Manor, New 35
York. Suitable non-woven materials include for example, a spun-bonded polypropylene commercially available from Reemay Inc. of Old Hickory, TN.

The treated membrane may then be further processed to remove the surfactant with the use of various low molecular weight alcohols. This is accomplished by soaking or submerging the membrane in a solution of, for

example, water, isopropyl alcohol, methanol and/or glycerin. During this step, the surfactant which was originally mixed in solution with the perfluoro ion exchange material is removed. Slight swelling of the membrane occurs. The perfluoro ion exchange material remains within the pores of the base material as it is not soluble in the lower molecular weight alcohol.

The membrane is further treated by boiling in a suitable swelling agent, preferably water causing it to then slightly swell in the x and y direction. Additional swelling occurs in the z-direction. A composite membrane results having a higher ion transport rate (at least 20 times higher than in its unswollen state) that is also strong. The swollen membrane retains its mechanical integrity unlike the membranes consisting only of the perfluoro ion exchange material and simultaneously maintains desired ionic transport characteristics. A correlation exists between the content of the swelling agent within the membrane structure and transport properties of the membrane. A swollen membrane will transport chemical species faster than an unswollen membrane.

Although the membrane has excellent long term chemical stability, it can be susceptible to poison by organics. The organics can be removed by regeneration in which the membrane is boiled in a strong acid such as nitric or chromic acid.

To prepare the inventive membrane, a support structure such as a polypropylene woven fabric may first be laminated to the untreated base membrane by any conventional technique including hot roll lamination, ultrasonic lamination, adhesive lamination, forced hot air lamination so long as the technique does not damage the integrity of the membrane. A solution is prepared containing a perfluorosulfonic acid resin in solvent mixed with one or more surfactants. The solution may be applied to the membrane by any conventional coating technique including forwarding roll coating, reverse roll coating, gravure coating, doctor coating, kiss coating, as well as dipping, brushing, painting, and spraying so long as the liquid solution is able to penetrate the interstices and micropores of the base material. Excess solution from the surface of the membrane may be removed. The treated membrane is then immediately placed into an oven to dry. Oven temperatures may range from 60-200°C, preferably 120-160°C so as to lock the perfluoro ion polymer inside the membrane and prevent it from migrating to the surface during drying. This step may be repeated until the membrane becomes completely transparent. Typically between 2 to 8 treatments are required but the actual number of treatments is dependent on the surfactant concentration and

thickness of the membrane. If the membrane is prepared without a support structure, both sides of the membrane may be treated simultaneously thereby reducing the number of treatments required.

5 The oven treated membrane is then soaked in a low molecular weight alcohol, as described above to remove the surfactant. The membrane is then boiled as described above in a swelling agent under pressure ranging from 0 to 20 atmospheres absolute thereby increasing the amount of swelling agent the treated membrane is capable of holding.

10 Alternatively, the ion exchange material may be applied to the membrane without the use of a surfactant. This procedure requires additional exposure to the perfluoro ion exchange resin but does not then need to be soaked in alcohol. A vacuum may also be used to draw the ion exchange material into the membrane.

15 Another alternative to the process of preparing the inventive membrane involves the selection and use of a surfactant having low water solubility with the perfluoro ion solution. Surfactants with low water solubility include Zonyl® FSO, a fluorocarbon based surfactant commercially available from E. I. DuPont de Nemours, Inc. By using this type of surfactant, the heat treatment step may be eliminated. The resulting ion exchange treated membrane made
20 by this process may be used for different aqueous applications and other chemical environments without any effect due to the surfactant.

25 Because the base membrane is exceptionally thin (at most 1 mil) (0.025 mm) with the resulting composite membrane being very thin and only marginally distorted in the x and y directions, it is able to selectively transport ions at a faster rate than heretofore has been achieved with only a slight lowering of the selectivity characteristics of the membrane.

The following testing procedures were employed on the samples prepared as described in the examples described below.

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TEST PROCEDURES

AIR PERMEABILITY - Gurley Number Method

35 The resistance of samples to air flow was measured by a Gurley densometer (ASTM D726-58) manufactured by W. & L. E. Gurley & Sons. The results were reported in terms of Gurley Number defined as the time in seconds for 100 cubic centimeters of air to pass through 1 square inch (6.45 sq. cm.) of a test sample at a pressure drop of 4.88 inches (12.4 cm.) of water.

Strength testing was carried out on an Instron Model 1122. Samples were one inch wide. Gauge length (distance between clamps) was two inches (5.08 cm.). Samples were pulled at a rate of 500% per minute. The cross head speed was 20 inches per minute.

Thickness of the base material was determined with the use of a snap gauge (Johannes Kafer Co. Model No. F1000/302). Measurements were taken in at least four areas of each specimen. Thickness of the dried composite membrane were also obtained with the use of the snap gauge. Thicknesses of swollen samples were not measurable due to the compression or residual water on the surface of the swollen membrane with the snap gauge. Thickness measurements of the swollen membrane were also not able to be obtained with the use of scanning electron microscopy due to interferences with the swelling agents.

Potassium Acetate Method

Moisture vapor transmission rates were determined by the following procedure. Approximately 70 ml. of a solution consisting of 35 parts by weight of potassium acetate and 15 parts by weight of distilled water was placed into a 133 ml. polypropylene cup, having an inside diameter of 6.5 cm. at its mouth. An expanded polytetrafluoroethylene (PTFE) membrane having a minimum MVTR of approximately 85,000 g/m²-24 hr as tested by the method described in U.S. Patent 4,862,730 to Crosby and available from W. L. Gore & Associates, Inc. of Newark, Delaware, was heat sealed to the lip of the cup to create a taut, leakproof, microporous barrier containing the solution.

30 A similar expanded PTFE membrane was mounted to the surface of a water bath. The water bath assembly was controlled at 23°C plus or minus 0.2°C, utilizing a temperature controlled room and a water circulating bath.

The sample to be tested was allowed to condition at a temperature of 23°C and a relative humidity of 50% prior to performing the test procedure. Samples were placed so the microporous polymeric membrane was in contact with the expanded polytetrafluoroethylene membrane mounted to the surface of the water bath and allowed to equilibrate for at least 15 minutes prior to the introduction of the cup assembly.

Water transport was provided by the driving force between the water in the water bath and the saturated salt solution providing water flux by diffusion in that direction. The sample was tested for 10 minutes and the cup assembly was then removed, weighed again within 1/1000g.

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Peel strength or membrane adhesion strength tests were conducted on samples prepared with reinforced backings. Test samples were prepared having dimensions of 3 inch by 3.5 inch (7.62 cm x 8.89 cm). A 4 inch by 4 inch (10.2 cm x 10.2 cm) chrome steel plate with an Instron tensile test machine Model No. 1000 were also used. Double coated vinyl tape (3M - #419) 1 inch (2.54 cm) wide was placed over the edges of the chrome plate so that tape covered all edges of the plate.

The plate and sample were then installed in the Instron in a horizontal position. The upper crosshead was lowered so that the jaws of the machine closed flat and tightened upon the sample. The upper crosshead was then slowly raised pulling the layer of perfluoro ion material from the reinforced backing. When the composite membrane had been detached from the reinforced backing, the test was complete. Adhesion strength was estimated from the average strength needed to pull the composite membrane from the reinforced backing.

The electrical conductance in the Z-direction (otherwise known as through conductance) was measured. A sample of swollen composite membrane (cut to a 1 inch diameter circle) was placed between two 0.680 inch (1.73 cm) diameter copper contacts. A 5 lb. (2268 g) weight was placed above the top contact. The contacts were connected to a Hewlett Packard Model 3478A multimeter. The resistance was then read. Prior to this measurement, the thickness of the dried preswollen composite membrane was

determined as described above. Conductance was calculated according to the formula:

$$C = 1/R$$

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wherein R = resistance measured in ohms

C = conductance measured in mhos

IONIC CONDUCTION RATE

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The ~~conductivity of the~~ composite membrane was tested to measure the ionic conduction rate in terms of micromhos per minute. This test was performed with two 900 ml. compartments between which the treated membrane was placed. The exposed surface area of membrane was 7.07 sq in. (45.65 sq cm). One compartment (the retentate side) was filled with 1 M NaCl solution. The other side (the permeate side) was filled with pure distilled water. Both compartments were stirred continuously and at the same speed with two electric mixers using polypropylene impellers. The ~~conductivity of the~~ ^{conductance} permeate side was recorded every 5 minutes for an hour with a hand-held conductivity meter, Omega Model No. PHH80. The total ionic conduction rate was determined by taking the average slope of a graph of ~~conductivity~~ ^{conductance} over time.

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LINEAR EXPANSION TESTS

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The percentage swelling in the x- and y-directions were determined. The length and width of the composite membranes were first measured with a Mitutoyo Model # 505-627-50 caliper prior to boiling and swelling. Final measurements were taken after the samples were boiled and swelled. Percent linear expansion were then calculated for both the x- and y-directions.

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WEIGHT CHANGE

The percent weight change on samples was also prepared. Here composite membranes were weighed prior to boiling and swelling and then after swelling. All weight measurements were taken with Mettler Balance, Model No. AT400. The percent weight change was then calculated.

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Example 1:

A sample of expanded polytetrafluoroethylene membrane made in accordance with the teachings of U.S. Patent No. 3,593,566, herein incorporated by reference. The membrane, with a nominal thickness of 0.75
5 mils (0.02 mm) and a 0.2 micrometer pore size, was mounted on a 6 inch wooden embroidery hoop. A solution was prepared comprising 95% by volume of a perfluorosulfonic acid/tetra-fluoroethylene copolymer resin (in H+ form) in a solution of low molecular weight alcohols comprising propanol, butanol, and methanol known as Nafion NR-50 (1100 EW) commercially
10 available from E.I. DuPont de Nemours, Inc. and 5% of a nonionic surfactant of octyl phenoxy poly ethoxyethanol known as Triton X100 commercially available from Rohm & Haas of Philadelphia, PA. This solution was brushed on both sides of the membrane so as to impregnate and substantially occlude the micropore structure. The sample was then dried in the oven at 140°C for
15 30 seconds. The procedure was repeated two more times to fully occlude the micropores. The sample was then soaked in isopropanol for 5 minutes to remove the surfactant. After rinsing with distilled water and allowing to dry at room temperature, a final coat of the Nafion-surfactant solution as described above was applied. The wet membrane was again dried in the oven at 140°C
20 for 30 seconds and soaked in isopropanol for 2 minutes. The membrane was finally boiled in distilled water for 10 minutes under atmospheric pressure to swell the treated membrane. Gurley numbers for this material are summarized in Table 3. Ionic conductive rates are summarized in Table 4. The strength modulus may be found in Table 5; percent linear expansion may be found in
25 Table 6; and percent weight change of this sample may be found in Table 7. The swollen membrane was later dried to a dehydrolyzed state in an oven at 140°C for 30 seconds. The thickness of the dried composite membrane was measured and found to be approximately the same thickness as the base material.

Example 2

A sample of expanded porous PTFE membrane made in accordance with the teachings of U.S. Patent No. 3,593,566 having a pore size of 0.2 micrometers and nominal thickness of 0.75 mils (0.02 mm) and a Gurley
35 Densometer air flow of 2-4 seconds was placed on top of a netting of polypropylene obtained from Conwed Plastics Corp. of Minneapolis, MN. The two materials were bonded together on a laminator with 10 psig pressure, a speed of 15 feet per minute and a temperature of 200°C. No adhesives were

used. The reinforced membrane sample was then placed on a 6 inch wooden embroidery hoop. A solution of 96% by volume of a perfluorosulfonic acid TFE copolymer resin in alcohol Nafion NR-50 (1100 EW) commercially available from E.I. DuPont de Nemours, Inc. and 4% of the nonionic surfactant Triton X-100 obtained from Rohm & Haas was prepared. This solution was brushed on the membrane side only to substantially occlude the micropores and the sample was dried in an oven at 130°C. This procedure was repeated three more times to fully occlude the micropores. The sample was then baked in an oven at 140°C for 5 minutes. The sample was soaked in isopropanol for 5 minutes to remove the surfactant. The membrane was then boiled in distilled water for 30 minutes under atmospheric pressure causing the treated membrane to swell. Gurley numbers for this material are summarized in Table 3.

This sample was tested for its peel strength in accordance with the method described above. The linear bond strength was found to be 2.06 lb./sq. in. (1450 kg/m²).

Example 3

A sample of expanded porous polytetrafluoroethylene membrane made in accordance with the teachings of U.S. Patent No. 3,593,566, having a thickness of 0.5 mils (0.01 mm) with a pore size of 0.2 micrometer was mounted on a 6 inch wooden embroidery hoop. A solution of 100% Nafion resin solution, perfluorosulfonic acid/TFE copolymer resin in a solvent mixture of propanol, butanol, and methanol known commercially from E. I. DuPont de Nemours, Inc. as Nafion® solution NR-50 (1100 EW) without the addition of any surfactants was brushed onto both sides of the membrane to substantially occlude the micropores. The sample was then placed in an oven at 160°C to dry. This procedure was repeated four more times until the membrane was completely transparent and the micropores were fully occluded. The sample was then boiled in distilled water for 30 minutes at atmospheric pressure causing the membrane to swell. Gurley numbers for this material are summarized in Table 3. The electrical conductivity was measured and summarized in Table 2.

Example 4

A sample of expanded porous polytetrafluoroethylene membrane made in accordance with the teachings of U.S. Patent No. 3,593,566 having a thickness of 0.5 mils (0.01 mm) and a pore size 0.2 micrometers was mounted

on a 6 inch wooden embroidery hoop. A solution of 99% by volume Nafion NR-50 commercially available from E.I. DuPont de Nemours, Inc. and 1% surfactant mixture was prepared. The surfactant mixture consisted of 50% of a nonionic surfactant, Triton X-100 commercially available from Rohm & Haas Corp. and 50% Zonyl FSO commercially available from E.I. DuPont de Nemours, Inc. This solution was brushed on both sides of the membrane and was allowed to dry at room temperature. This procedure was repeated 4 more times until the sample was completely transparent and to fully occlude the micropores. The sample was not treated so as to remove the surfactant. The composite membrane was boiled in distilled water for 5 minutes causing the membrane to swell. The Gurley number for this material is summarized in Table 3.

Example 5

A sample of expanded porous polytetrafluoroethylene membrane made in accordance with the teachings of U.S. Patent No. 3,593,566, having a thickness of 0.5 mils (0.01 mm) with a pore size of 0.2 micrometer was mounted onto a 6 inch wooden embroidery hoop. A solution of 95% by volume Nafion NR-50 (1100 EW) commercially available from E.I. DuPont de Nemours, Inc. and 5% of a nonionic surfactant, Triton X-100 commercially available from Rohm & Haas was prepared. The solution was brushed on both sides of the membrane with a foam brush and the excess was wiped off. The wet membrane was dried in an oven at 140°C for 30 seconds. Three additional coats of solution were applied to the membrane in the same manner to fully occlude the micropores. The membrane was then soaked in isopropanol for 2 minutes to remove the surfactant. The membrane was rinsed with distilled water and allowed to dry at room temperature. A final treatment of the Nafion-Triton solution was applied. The wet membrane was dried in the oven at 140°C for 30 seconds, then soaked in isopropanol for 2 minutes. Finally, the membrane was boiled in distilled water for 5 minutes. Moisture vapor transmission rates for the treated membrane were measured and are summarized in Table 1. The Gurley number of the treated membrane are summarized in Table 3.

Example 6

A sample of expanded porous polytetrafluoroethylene membrane made in accordance with the teachings of U.S. Patent No. 53,593,566, having a nominal thickness of 0.75 mils (0.02 mm) and a pore size of 0.2 micrometers

was mounted onto a 6 inch wooden embroidery hoop. The Gurley Densometer air flow on this membrane was 2-4 seconds. A solution of 95% by volume Nafion NR-50 (1100 EW) commercially available from E.I. DuPont de Nemours, Inc. and 5% Triton X-100 non-ionic surfactant from Rohm & Haas was prepared. The solution was brushed on both sides of the membrane with a foam brush and the excess was wiped off. The wet membrane was dried in the oven at 140°C for 30 seconds. Three additional coats of solution were applied in the same manner. The membrane was then soaked in isopropanol for 2 minutes. After rinsing with distilled water and allowing to dry at room temperature, a final coat of the Nafion surfactant solution was applied. The wet membrane was dried in the oven at 140°C for 30 seconds, then soaked in isopropanol for 2 minutes. This material was not boiled. No swelling other than the minor swelling during the surfactant removal occurred. The ionic conduction rate for this material is presented in Table 4.

Example 7

A sample of expanded porous polytetrafluoroethylene membrane made in accordance with the teachings of U.S. Patent No. 3,593,566, having a nominal thickness of 0.75 mils (0.02 mm) and a pore size of 0.2 micrometers was mounted onto a 5 inch plastic embroidery hoop. The Gurley Densometer air flow on this membrane was 2-4 seconds. A solution of 95% by volume Nafion NR-50 (1100 EW) commercially available from E.I. DuPont de Nemours, Inc. and 5% Triton X-100 non-ionic surfactant from Rohm & Haas was prepared. The solution was brushed on both sides of the membrane with a foam brush and the excess was wiped off. The wet membrane was dried in the oven at 140°C for 30 seconds. Two additional coats of solution were applied in the same manner so as to fully occlude the micropores. The membrane was then soaked in isopropanol for 2 minutes. After rinsing with distilled water and allowing to dry at room temperature, a final coat of the same Nafion NR-50 Triton X-100 solution was applied. The wet membrane was dried in the oven at 140°C for 30 seconds and then soaked in isopropanol for 2 minutes to remove the surfactant. The sample was rinsed and dried at room temperature. No boiling occurred.

This sample was weighed before it was mounted on the 5 inch plastic hoop. Following treatment, it was removed from the hoop and weighed again. The ion exchange polymer content was directly calculated by determining the weight change before and after treatment. The ion exchange content for this sample was found to be 98.4 mg or 9.81 grams per square meter of

membrane. A sample of Nafion 115 (5 mils, 0.13 mm) commercially available from E. I. DuPont de Nemours, Inc. was cut to a 1 inch (25.4 mm) by 1 inch (25.4 mm) sample, weighed and found to be 216 grams per square meter.

Nafion Comparative Samples

Nafion 117, a perfluorosulfonic acid cation exchange membrane, unreinforced film of 1100 equivalent weight commercially available from E. I. DuPont de Nemours Co., Inc., having a quoted nominal thickness of 7 mils (0.18 mm) was obtained. The samples, originally in the hydrated swollen state were measured in the x- and y-directions and weighed. The samples were then dried at room temperature to an unswollen state and then remeasured from which expansion and weight change measurements found in Tables 6 and 7 were calculated. Nafion 115, a perfluorosulfonic acid cation exchange membrane, unreinforced film of 1100 equivalent weight also commercially available from E. I. DuPont de Nemours, Inc., having a nominal thickness of 5 mils (0.1 mm) was obtained. This sample was also obtained commercially in the hydrated swollen state.

TABLE 1

Moisture Vapor Transmission Rates (MVTR)

| Sample ID* | MVTR (grams/m ² -24 hrs. |
|------------|-------------------------------------|
| 5 | 25,040 |
| Nafion 117 | 23,608 |

*Measurements were obtained on samples in their swollen state.

TABLE 2

Electrical Conductivity

| Sample ID* | Conductance (micromhos) |
|------------|-------------------------|
| 3 | 1,277 |
| Nafion 117 | 1,214 |

*Measurements were obtained on samples in their swollen state.

TABLE 3
Gurley Numbers

| Sample ID | Thickness (mm)* | Base Material Gurley No. (sec) | Final Swollen Membrane Gurley Number (sec) |
|-----------|-----------------|--------------------------------|--|
| 1 | 0.02 | 2-4 | Total occlusion |
| 2 | 0.02 | 2-4 | Total occlusion |
| 3 | 0.01 | 2-4 | Total occlusion |
| 4 | 0.01 | 2-4 | Total occlusion |
| 5 | 0.01 | 2-4 | Total occlusion |

5 *Thickness measurements were obtained on samples prior to swelling in dried state.

TABLE 4
Ionic Conduction Rate

| Sample ID | Ionic Conduction Rate (micromhos/minute) |
|------------|--|
| 1 | 119 (swollen) |
| 6 | 5.1 (unswollen) |
| Nafion 115 | 15.9 (swollen) |

TABLE 5
Strength Modulus

| Sample ID | Thickness (mm) | Strength Modulus (lb per dry sq. in.)** |
|------------|----------------|---|
| 1 | 0.02* | 15150 |
| Nafion 115 | 0.13 | 12750 |

*Thickness measurements were obtained prior to swelling in dried state.

15 **Strength modulus measurements were obtained in the swollen state.

TABLE 6
Percent Linear Expansion

| Sample ID | Unswollen (x) (mm) | Swollen (x) (mm) | Average % Expansion in x-direction | Unswollen (y) (mm) | Swollen (y) (mm) | Average % Expansion in y-direction |
|------------|--------------------------|------------------------|--|--------------------------|------------------------|--|
| 1 | 124.4 | 124.4 | -- | 123.3 | 123.3 | -- |
| Nafion 117 | 125.5 | 137.7 | +9.7 | 127.3 | 149.7 | +17 |

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TABLE 7
Percent Weight Change

| Sample ID | Unswollen wt (g) | Swollen wt (wet) (g) | % weight change (wet) |
|------------|---------------------|-------------------------|--------------------------|
| 1 | 0.2515 | 1.0273 | +308% |
| Nafion 117 | 5.5700 | 7.5106 | +35% |